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N-(3-Chloro-4-fluorophenyl)-2-(naphthalen-1-yl)acetamide

A. S. Praveen,[a] Jerry P. Jasinski,[b]* James A. Golen,[b] B. Narayana[c] and H. S. Yathirajan[a]

[a]Department of Studies in Chemistry, University of Mysore, Manasagangotri, Mysore 570 006, India, [b]Department of Chemistry, Keene State College, 229 Main Street, Keene, NH 03435-2001, USA, and [c]Department of Studies in Chemistry, Mangalore University, Mangalagangotri 574 199, India
Correspondence e-mail: jjasinski@keene.edu

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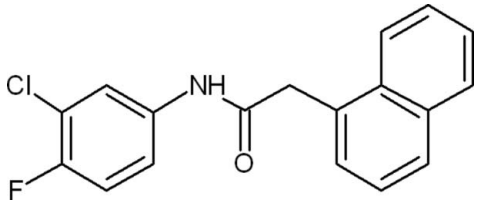
Key indicators: single-crystal X-ray study; T = 173 K; mean σ(C–C) = 0.003 Å; R factor = 0.044; wR factor = 0.119; data-to-parameter ratio = 18.2.

In the title compound, $C_{18}H_{13}ClFNO$, the dihedral angle between the mean planes of the chloro- and fluoro-substituted benzene ring and the naphthalene ring system is 60.5 (8)°. In the crystal, molecules are linked by N—H···O hydrogen bonds, forming a zigzag chain along [101].

Related literature

For the structural similarity of *N*-substituted 2-arylacetamides to the lateral chain of natural benzylpenicillin, see: Mijin & Marinkovic (2006); Mijin *et al.* (2008). For the coordination abilities of amides, see: Wu *et al.* (2008, 2010). For related structures, see: Davis & Healy (2010); Li *et al.* (2010); Li & Wu (2010); Wang *et al.* (2010); Xiao *et al.* (2010). For standard bond lengths, see: Allen *et al.* (1987).



Experimental

Crystal data

$C_{18}H_{13}ClFNO$
M_r = 313.74
Monoclinic, $P2_1/n$
a = 8.096 (6) Å
b = 23.323 (6) Å
c = 8.404 (3) Å
β = 110.83 (5)°
V = 1483.4 (13) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 0.27 mm^{-1}
T = 173 K
0.30 × 0.18 × 0.10 mm

Data collection

Oxford Diffraction Oxford Xcalibur Eos Gemini diffractometer
Absorption correction: multi-scan (*CrysAlis RED*; Oxford Diffraction, 2010)
T_{min} = 0.924, T_{max} = 0.974
13979 measured reflections
3679 independent reflections
2947 reflections with $I > 2\sigma(I)$
R_{int} = 0.024

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.044
$wR(F^2)$ = 0.119
S = 1.04
3679 reflections
202 parameters
1 restraint
H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{max}$ = 0.25 e Å^{-3}
$\Delta\rho_{min}$ = −0.31 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
N1—H1*N*···O1^{i}	0.85 (1)	2.12 (2)	2.914 (2)	157 (2)

Symmetry code: (i) $x+\frac{1}{2}, -y+\frac{1}{2}, z+\frac{1}{2}$.

Data collection: *CrysAlis PRO* (Oxford Diffraction, 2010); cell refinement: *CrysAlis PRO*; data reduction: *CrysAlis RED* (Oxford Diffraction, 2010); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: IS2737).

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supplementary materials

Acta Cryst. (2011). E**67**, o1826 [**doi:10.1107/S1600536811024597**]

N-(3-Chloro-4-fluorophenyl)-2-(naphthalen-1-yl)acetamide

A. S. Praveen, J. P. Jasinski, J. A. Golen, B. Narayana and H. S. Yathirajan

Comment

N-Substituted 2-arylacetamides are very interesting compounds because of their structural similarity to the lateral chain of natural benzylpenicillin (Mijin *et al.*, 2006, 2008). Amides are also used as ligands due to their excellent coordination abilities (Wu *et al.*, 2008, 2010). Crystal structures of some acetamide derivatives, viz., 2-(4-bromophenyl)-N-(2-methoxyphenyl)acetamide (Xiao *et al.*, 2010), N-benzyl-2-(3-chloro-4-hydroxyphenyl)acetamide (Davis & Healy, 2010), 2-(2,2-dimethyl-2,3-dihydro-1-benzofuran-7-yloxy)-N-(o-tolyl)acetamide (Li *et al.*, 2010), N-benzyl-2-(2-bromophenyl)-2-(2-nitrophenoxy) acetamide (Li & Wu, 2010) and N-(4-chlorophenyl)-2-(8-quinolyloxy)acetamide monohydrate (Wang *et al.*, 2010) have been reported. In view of the importance of amides, we report herein the crystal structure of the title compound, (I), $C_{18}H_{13}ClFNO$.

In the title compound, $C_{18}H_{13}ClFNO$, the dihedral angle between the mean planes of the chloro, fluoro substituted benzene ring and the naphthalene-1-yl ring is 60.5 (8)° (Fig. 2). Bond distances are in normal ranges (Allen *et al.*, 1987). Crystal packing is stabilized by N—H···O hydrogen bonds (Fig. 3 and Table 1).

Experimental

Naphthalen-1-ylacetyl chloride (0.204 g, 1 mmol) and 3-chloro-4-fluoroaniline (0.145 g, 1 mmol) were dissolved in dichloromethane (20 mL). The mixture was stirred in presence of triethylamine at 273 K for about 3 h (Fig. 1). The contents were poured into 100 ml of ice-cold aqueous hydrochloric acid with stirring, which was extracted thrice with dichloromethane. Organic layer was washed with saturated $NaHCO_3$ solution and brine solution, dried and concentrated under reduced pressure to give the title compound (I). Single crystals were grown from toluene by the slow evaporation method (M.P.: 421 K).

Refinement

The N-bound H atom was located in a difference Fourier map and refined isotropically with a distance restraint of N—H = 0.86 (2) Å. All of the remaining H atoms were placed in their calculated positions and then refined using the riding model, with C—H lengths of 0.95 Å (CH) or 0.99 Å (CH_2). Isotropic displacement parameters for these atoms were set to 1.19-1.21 (CH) or 1.20 (CH_2) times U_{eq} of the parent atom.

Figures

Fig. 1. Reaction scheme of the title compound, (I).

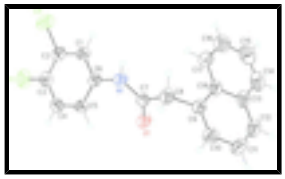

Fig. 2. Molecular structure of the title compound, showing the atom labeling scheme and 50% probability displacement ellipsoids.

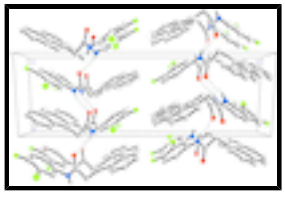

Fig. 3. Packing diagram of the title compound viewed down the *c* axis. Dashed lines represent N—H···O hydrogen bonds.

N-(3-Chloro-4-fluorophenyl)-2-(naphthalen-1-yl)acetamide

Crystal data

$C_{18}H_{13}ClFNO$	$F(000) = 648$
$M_r = 313.74$	$D_x = 1.405$ Mg m^{-3}
Monoclinic, $P2_1/n$	Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
Hall symbol: -P 2yn	Cell parameters from 4655 reflections
$a = 8.096$ (6) Å	$\theta = 3.1$–$32.5°$
$b = 23.323$ (6) Å	$\mu = 0.27$ mm^{-1}
$c = 8.404$ (3) Å	$T = 173$ K
$\beta = 110.83$ (5)°	Block, colorless
$V = 1483.4$ (13) Å^3	0.30 × 0.18 × 0.10 mm
$Z = 4$	

Data collection

Oxford Diffraction Oxford Xcalibur Eos Gemini diffractometer	3679 independent reflections
Radiation source: Enhance (Mo) X-ray Source	2947 reflections with $I > 2\sigma(I)$
graphite	$R_{int} = 0.024$
Detector resolution: 16.1500 pixels mm^{-1}	$\theta_{max} = 28.3°$, $\theta_{min} = 3.1°$
ω scans	$h = -10 \rightarrow 10$
Absorption correction: multi-scan (*CrysAlis RED*; Oxford Diffraction, 2010)	$k = -31 \rightarrow 31$
$T_{min} = 0.924$, $T_{max} = 0.974$	$l = -10 \rightarrow 11$
13979 measured reflections	

Refinement

Refinement on F^2	Primary atom site location: structure-invariant direct methods
Least-squares matrix: full	Secondary atom site location: difference Fourier map
$R[F^2 > 2\sigma(F^2)] = 0.044$	Hydrogen site location: inferred from neighbouring sites
$wR(F^2) = 0.119$	H atoms treated by a mixture of independent and constrained refinement

S = 1.04	$w = 1/[\sigma^2(F_o^2) + (0.0495P)^2 + 0.4725P]$ where $P = (F_o^2 + 2F_c^2)/3$
3679 reflections	$(\Delta/\sigma)_{max} = 0.001$
202 parameters	$\Delta\rho_{max} = 0.25$ e Å^{-3}
1 restraint	$\Delta\rho_{min} = -0.31$ e Å^{-3}

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on F^2, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating *R*-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
Cl1	0.38936 (7)	0.11723 (2)	1.00561 (6)	0.06507 (18)
F1	0.17103 (16)	0.03432 (5)	0.75440 (15)	0.0629 (3)
O1	0.21705 (14)	0.22471 (5)	0.22432 (14)	0.0486 (3)
N1	0.43807 (16)	0.20845 (6)	0.47811 (16)	0.0379 (3)
H1N	0.538 (2)	0.2204 (7)	0.544 (2)	0.046*
C1	0.4025 (2)	0.16306 (7)	0.71760 (19)	0.0380 (3)
H1B	0.4762	0.1917	0.7880	0.046*
C2	0.3379 (2)	0.11924 (7)	0.7882 (2)	0.0414 (4)
C3	0.2317 (2)	0.07766 (7)	0.6847 (2)	0.0428 (4)
C4	0.1869 (2)	0.07943 (7)	0.5115 (2)	0.0426 (4)
H4A	0.1126	0.0507	0.4417	0.051*
C5	0.2506 (2)	0.12309 (6)	0.43962 (19)	0.0385 (3)
H5A	0.2205	0.1247	0.3197	0.046*
C6	0.35908 (18)	0.16496 (6)	0.54278 (18)	0.0346 (3)
C7	0.36895 (18)	0.23348 (6)	0.32430 (18)	0.0347 (3)
C8	0.4966 (2)	0.27172 (7)	0.2798 (2)	0.0407 (3)
H8A	0.5850	0.2870	0.3858	0.049*
H8B	0.5604	0.2488	0.2209	0.049*
C9	0.40433 (19)	0.32093 (6)	0.16728 (19)	0.0373 (3)
C10	0.3683 (2)	0.31871 (8)	−0.0039 (2)	0.0460 (4)
H10A	0.4071	0.2866	−0.0507	0.055*
C11	0.2745 (2)	0.36315 (9)	−0.1131 (2)	0.0555 (5)
H11A	0.2509	0.3606	−0.2320	0.067*
C12	0.2185 (2)	0.40899 (8)	−0.0499 (2)	0.0545 (5)
H12A	0.1538	0.4383	−0.1248	0.065*
C13	0.2547 (2)	0.41419 (7)	0.1263 (2)	0.0449 (4)
C14	0.2016 (3)	0.46210 (8)	0.1974 (3)	0.0609 (5)

H14A	0.1363	0.4918	0.1247	0.073*
C15	0.2418 (3)	0.46677 (9)	0.3675 (3)	0.0695 (6)
H15A	0.2056	0.4997	0.4130	0.083*
C16	0.3361 (3)	0.42340 (9)	0.4765 (3)	0.0633 (5)
H16A	0.3648	0.4274	0.5958	0.076*
C17	0.3873 (2)	0.37576 (7)	0.4146 (2)	0.0482 (4)
H17A	0.4493	0.3463	0.4906	0.058*
C18	0.34938 (19)	0.36936 (7)	0.2373 (2)	0.0384 (3)

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cl1	0.0760 (4)	0.0795 (4)	0.0392 (2)	−0.0043 (3)	0.0199 (2)	0.0062 (2)
F1	0.0707 (7)	0.0521 (6)	0.0693 (7)	−0.0078 (5)	0.0291 (6)	0.0142 (5)
O1	0.0360 (6)	0.0473 (6)	0.0456 (6)	−0.0067 (5)	−0.0064 (5)	0.0111 (5)
N1	0.0282 (6)	0.0420 (7)	0.0344 (6)	−0.0039 (5)	−0.0002 (5)	0.0035 (5)
C1	0.0342 (7)	0.0390 (8)	0.0366 (7)	0.0023 (6)	0.0072 (6)	−0.0021 (6)
C2	0.0395 (8)	0.0478 (9)	0.0359 (7)	0.0078 (7)	0.0120 (6)	0.0047 (6)
C3	0.0392 (8)	0.0380 (8)	0.0522 (9)	0.0042 (6)	0.0175 (7)	0.0077 (7)
C4	0.0385 (8)	0.0351 (7)	0.0492 (9)	−0.0008 (6)	0.0096 (7)	−0.0033 (6)
C5	0.0363 (8)	0.0395 (8)	0.0353 (7)	0.0004 (6)	0.0072 (6)	−0.0018 (6)
C6	0.0273 (6)	0.0363 (7)	0.0358 (7)	0.0044 (5)	0.0060 (6)	0.0029 (6)
C7	0.0307 (7)	0.0316 (7)	0.0352 (7)	0.0031 (5)	0.0037 (6)	0.0001 (5)
C8	0.0315 (7)	0.0425 (8)	0.0433 (8)	0.0013 (6)	0.0073 (6)	0.0056 (6)
C9	0.0296 (7)	0.0401 (8)	0.0374 (7)	−0.0049 (6)	0.0058 (6)	0.0060 (6)
C10	0.0420 (9)	0.0526 (10)	0.0402 (8)	−0.0100 (7)	0.0107 (7)	−0.0010 (7)
C11	0.0505 (10)	0.0740 (13)	0.0348 (8)	−0.0145 (9)	0.0064 (7)	0.0117 (8)
C12	0.0429 (9)	0.0570 (11)	0.0527 (10)	−0.0052 (8)	0.0037 (8)	0.0241 (8)
C13	0.0333 (8)	0.0420 (8)	0.0544 (9)	−0.0035 (6)	0.0094 (7)	0.0135 (7)
C14	0.0519 (11)	0.0419 (9)	0.0868 (15)	0.0050 (8)	0.0222 (10)	0.0154 (9)
C15	0.0762 (14)	0.0506 (11)	0.0893 (16)	0.0019 (10)	0.0386 (13)	−0.0080 (11)
C16	0.0754 (14)	0.0589 (12)	0.0604 (12)	−0.0043 (10)	0.0302 (11)	−0.0072 (9)
C17	0.0509 (10)	0.0481 (9)	0.0428 (9)	−0.0026 (7)	0.0133 (8)	0.0034 (7)
C18	0.0309 (7)	0.0390 (8)	0.0414 (8)	−0.0050 (6)	0.0081 (6)	0.0063 (6)

Geometric parameters (Å, °)

Cl1—C2	1.7242 (17)	C9—C10	1.363 (2)
F1—C3	1.3454 (19)	C9—C18	1.416 (2)
O1—C7	1.235 (2)	C10—C11	1.414 (3)
N1—C7	1.3457 (19)	C10—H10A	0.9500
N1—C6	1.407 (2)	C11—C12	1.343 (3)
N1—H1N	0.849 (14)	C11—H11A	0.9500
C1—C2	1.375 (2)	C12—C13	1.409 (3)
C1—C6	1.385 (2)	C12—H12A	0.9500
C1—H1B	0.9500	C13—C14	1.405 (3)
C2—C3	1.380 (2)	C13—C18	1.429 (2)
C3—C4	1.370 (2)	C14—C15	1.353 (3)
C4—C5	1.375 (2)	C14—H14A	0.9500

C4—H4A	0.9500	C15—C16	1.396 (3)
C5—C6	1.391 (2)	C15—H15A	0.9500
C5—H5A	0.9500	C16—C17	1.353 (3)
C7—C8	1.510 (2)	C16—H16A	0.9500
C8—C9	1.506 (2)	C17—C18	1.418 (2)
C8—H8A	0.9900	C17—H17A	0.9500
C8—H8B	0.9900		
C7—N1—C6	126.36 (13)	C10—C9—C18	119.18 (14)
C7—N1—H1N	117.4 (12)	C10—C9—C8	120.38 (15)
C6—N1—H1N	116.2 (12)	C18—C9—C8	120.41 (14)
C2—C1—C6	119.32 (14)	C9—C10—C11	121.43 (18)
C2—C1—H1B	120.3	C9—C10—H10A	119.3
C6—C1—H1B	120.3	C11—C10—H10A	119.3
C1—C2—C3	119.85 (15)	C12—C11—C10	120.37 (17)
C1—C2—Cl1	119.47 (13)	C12—C11—H11A	119.8
C3—C2—Cl1	120.68 (13)	C10—C11—H11A	119.8
F1—C3—C4	119.05 (15)	C11—C12—C13	120.77 (16)
F1—C3—C2	119.71 (15)	C11—C12—H12A	119.6
C4—C3—C2	121.24 (15)	C13—C12—H12A	119.6
C3—C4—C5	119.37 (15)	C14—C13—C12	122.34 (17)
C3—C4—H4A	120.3	C14—C13—C18	118.60 (17)
C5—C4—H4A	120.3	C12—C13—C18	119.06 (17)
C4—C5—C6	119.90 (14)	C15—C14—C13	121.24 (18)
C4—C5—H5A	120.0	C15—C14—H14A	119.4
C6—C5—H5A	120.0	C13—C14—H14A	119.4
C1—C6—C5	120.32 (14)	C14—C15—C16	120.3 (2)
C1—C6—N1	117.00 (13)	C14—C15—H15A	119.8
C5—C6—N1	122.56 (14)	C16—C15—H15A	119.8
O1—C7—N1	123.71 (15)	C17—C16—C15	120.9 (2)
O1—C7—C8	122.22 (14)	C17—C16—H16A	119.5
N1—C7—C8	114.01 (13)	C15—C16—H16A	119.5
C9—C8—C7	112.03 (13)	C16—C17—C18	120.59 (17)
C9—C8—H8A	109.2	C16—C17—H17A	119.7
C7—C8—H8A	109.2	C18—C17—H17A	119.7
C9—C8—H8B	109.2	C9—C18—C17	122.51 (14)
C7—C8—H8B	109.2	C9—C18—C13	119.17 (15)
H8A—C8—H8B	107.9	C17—C18—C13	118.31 (16)
C6—C1—C2—C3	0.3 (2)	C18—C9—C10—C11	−1.2 (2)
C6—C1—C2—Cl1	−179.17 (11)	C8—C9—C10—C11	176.99 (15)
C1—C2—C3—F1	178.93 (14)	C9—C10—C11—C12	0.1 (3)
Cl1—C2—C3—F1	−1.6 (2)	C10—C11—C12—C13	1.1 (3)
C1—C2—C3—C4	−0.8 (2)	C11—C12—C13—C14	178.56 (17)
Cl1—C2—C3—C4	178.65 (13)	C11—C12—C13—C18	−1.1 (2)
F1—C3—C4—C5	−179.02 (14)	C12—C13—C14—C15	−178.31 (18)
C2—C3—C4—C5	0.7 (2)	C18—C13—C14—C15	1.4 (3)
C3—C4—C5—C6	−0.1 (2)	C13—C14—C15—C16	−0.6 (3)
C2—C1—C6—C5	0.3 (2)	C14—C15—C16—C17	−0.8 (3)
C2—C1—C6—N1	−175.86 (14)	C15—C16—C17—C18	1.3 (3)

C4—C5—C6—C1	−0.4 (2)	C10—C9—C18—C17	−177.71 (15)
C4—C5—C6—N1	175.56 (14)	C8—C9—C18—C17	4.1 (2)
C7—N1—C6—C1	−150.76 (15)	C10—C9—C18—C13	1.2 (2)
C7—N1—C6—C5	33.2 (2)	C8—C9—C18—C13	−177.05 (13)
C6—N1—C7—O1	6.9 (3)	C16—C17—C18—C9	178.35 (17)
C6—N1—C7—C8	−170.51 (14)	C16—C17—C18—C13	−0.5 (2)
O1—C7—C8—C9	33.6 (2)	C14—C13—C18—C9	−179.71 (15)
N1—C7—C8—C9	−148.94 (14)	C12—C13—C18—C9	0.0 (2)
C7—C8—C9—C10	−100.15 (18)	C14—C13—C18—C17	−0.8 (2)
C7—C8—C9—C18	78.05 (18)	C12—C13—C18—C17	178.91 (15)

Hydrogen-bond geometry (Å, °)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
N1—H1N···O1^{i}	0.85 (1)	2.12 (2)	2.914 (2)	157.(2)

Symmetry codes: (i) *x*+1/2, −*y*+1/2, *z*+1/2.

Fig. 1

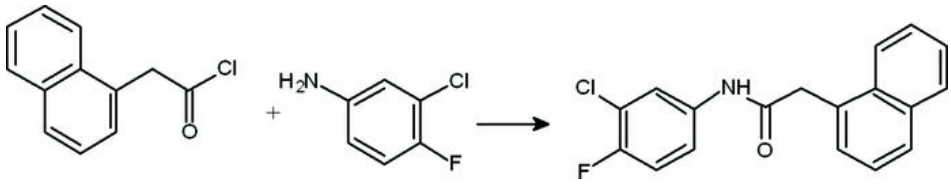

Fig. 2

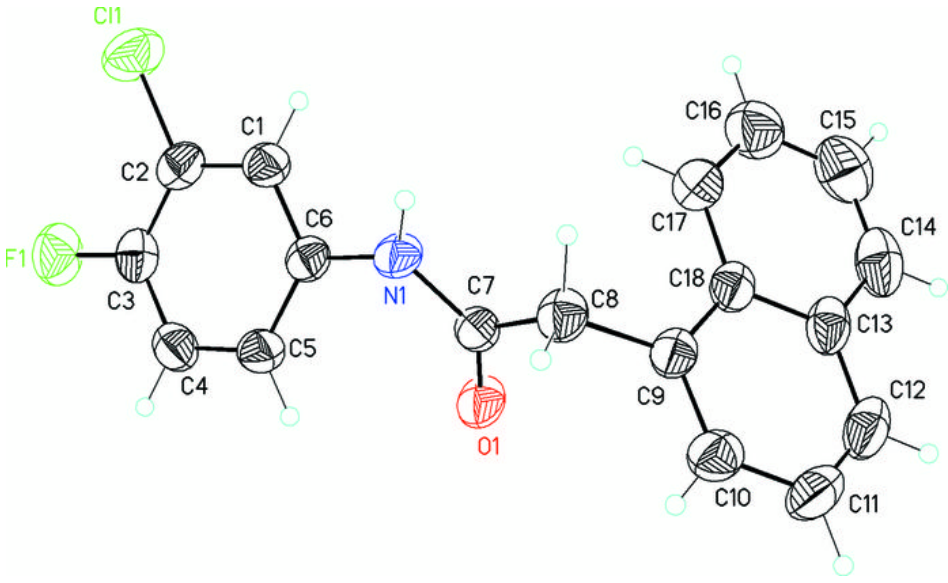
Cl1
C1
C2
C6
F1
C3
N1
C7
C8
C18
C16
C15
C17
C14
C13
C5
C4
C9
C12
O1
C10
C11

Fig. 3